Polyatomic Ions

Characteristics:

- Only compounds with a net charge (charge left over), mostly anions **(called oxyanions)**
- Most contain oxygen; ite and ate endings refer to the relative amount of oxygen
- You must memorize ions written in italics

Mercury I or Mercurous	Hg_2^{2+}
Ammonium	NH_4^+
Cyanide	CN^-
Cyanate	OCN^-
Carbonate	CO_3^{2-}
Hydrogen Carbonate (bicarbonate)	HCO_3^-
Acetate	$C_2H_3O_2^-$
Hypochlorite	ClO^-
Chlorite	ClO_2^-
Chlorate	ClO_3^-
Perchlorate	ClO_4^-
Chromate	CrO_4^{2-}
Dichromate	$Cr_2O_7^{2-}$
Permanganate	MnO_4^-
Nitrite	NO_2^-
Nitrate	NO_3^-
Hydroxide	OH^-
Peroxide	O_2^{2-}
Phosphite	PO_3^{3-}
Phosphate	PO_4^{3-}
Oxalate	$C_2O_4^{2-}$
Monohydrogen phosphate	HPO_4^{2-}
Dihydrogen phosphate	$H_2PO_4^-$
Sulfite	SO_3^{2-}
Sulfate	SO_4^{2-}
Hydrogen sulfite	HSO_3^-
Hydrogen Sulfate	HSO_4^-
Thiosulfate	$S_2O_3^{2-}$
Iodate	IO_3^-
Bromate	BrO_3^-
Perbromate	BrO_4^-
Hydrogen phosphate	HPO_4^{2}-
Silidate	SiO_3^{2-}
Hexafluorosilicate	SiF_6^{2-}
Selenate	SeO_4^{2-}
Tartrate	$C_4H_4O_6^{2-}$
Asenate	AsO_4^{3-}